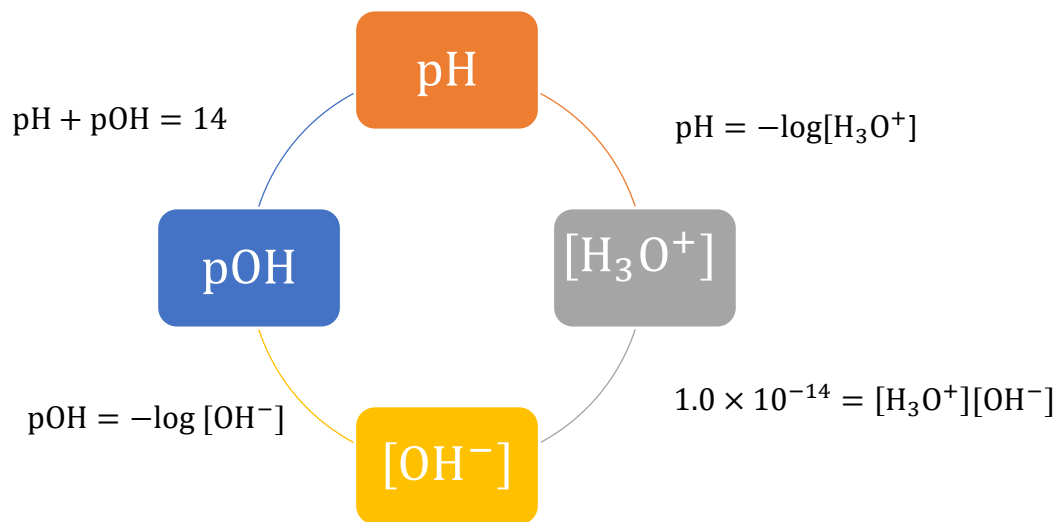


Test 3 Equations

Chapter 14



For weak acids

$$\% \text{ ionization} = \frac{[H_3O^+]_{eq}}{[HA]_0} \times 100$$

For weak bases

$$\% \text{ ionization} = \frac{[OH^-]_{eq}}{[B]_0} \times 100$$

$$1.0 \times 10^{-14} = K_a \times K_b$$

$$pK_a = -\log K_a$$

$$pK_b = -\log K_b$$